



HDH is a common medicament, a first generation drugs of antihistamines (H1 receptor antagonist). HDH H1 receptor antagonism is widely used in the treatment of pruritus, chronic and cholinergic urticaria and dermatoses. It has mild gastric secretion blockade, antibradykinin, anticholinergic, antiserotonergic, sedative, anxiolytic and anti-emetic properties [1-4]. The prescribed usage of HDH as a medicament does not have serious contra-indications. However, the geriatric patients are liable to suffer from secondary effects, such as sedation, cognitive impairment, confusion, blurred vision and urinary retention. A daily dosage of 50 to 100 mg is advisable for the treatment of tension and anxiety. An overdose of HDH or its combination with other sedatives, alcohols or lithium may be associated with hallucinations [5, 6].

The ability to design, construct and run an electroanalytical method that is more specific, sensitive and robust is crucial for drugs assay or quantification. Electroanalytical techniques are quite popular, not only to study chemical and biological species electrical behaviour, but also for quantitation purposes. To mention a few, it were used, for the detection of: resorcinol, a carbon nanotube paste electrode (CNTPE) modified with a graphite mixture [7]; dopamine, a CNTPE modified with poly basic blue B [8], a CPE modified with poly nicotinic acid [9] and poly phthalo blue [10]; indigo carmine, in food and water samples, a poly (glutamic acid) layered multi-walled (M-W) CNTPE [11]; tartrazine, a CPE modified with a polymer [12]; formoterol fumarate, a layered M-WCNTPE modified with poly(thiazole yellow-G) [13]; ciprofloxacin and estriol, an anionic surfactant (AS) and a CPE modified with a polymer [14]; riboflavin, AS, a polymer-modified electrode [14] and a C sensor modified with poly (DL-phenylalanine) [15]; curcumin, a CNT electrochemical sensor modified with a polymer [16]; and vanillin, in food samples, a poly(glutamic acid) functionalized M-WCNT and a graphite composite paste sensor [17].

Despite the electroanalytical techniques vast employability for the quantification of compounds of therapeutic interest, they are not widely used. By keeping features such as ease of performance, cost-effectiveness, and non-requirement of highly skilful operators, potentiometry with selective electroanalytical potentiometric sensors is highly useful for assaying therapeutic compounds such as HDH.

The extensive literature survey on the analytical methods reported for HDH quantification in pharmaceuticals revealed that there is a monography in United States Pharmacopoeia (USP) that adopted a high-performance liquid chromatography (HPLC) procedure, with uv-detection at 232 nm [18-23]. The other methods used different analytical techniques, such as gas [24], thin layer [25] and micellar liquid [26] chromatography, capillary zone electrophoresis [27], voltammetry [28], LC-MS (mass spectrometry) [29], potentiometry [30, 31], gravimetry [32], visible spectrophotometry [33-38] and titrimetry [33, 39-42]. The method reported by Sanrick and Janik [42] employed HDH precipitation with sodium tetraphenyl borate, and filtration and dissolution of the precipitate in acetone, followed by potentiometric titration with silver nitrate ( $\text{AgNO}_3$ ). The chromatographic, electrophoretic and voltammetry procedures for HDH detection are not simple and easy to apply. Despite of their high sensitivity, they are not cost effective, and need an operator with high skills. On the other end, methods such as

titrimetry and gravimetry require precipitation, filtration and titration of the residual precipitant. The required steps are complex, tedious and time-consuming. The spectrophotometric procedures involve the use of organic reagents, stringent experimental conditions and tedious extraction procedures.

In contrast to the limitations of the reported methods, this study presents a simple, novel and reliable potentiometric procedure for HDH determination, using the ion-pair complex of HDH and Orange II (ORG II) in a PVC matrix. A simple electronic device, and non-toxic, low cost and common standards and reagents were used in the development of this method, which has been optimized and validated. The procedure has showed good applicability to determine HDH in pharmaceuticals.

## **Experiment**

### ***Apparatus***

A digital dual channel potentiometer (PICO Chennai-32, India) and an Elico (Mumbai, India) pH-meter were used in the potential and pH recordings. Ag/AgCl and Al wire were used as WE and CE, respectively, to build an electrochemical cell for potential measurements.

### ***Reagents and solutions***

All chemicals used were of analytical reagent grade. Spectroscopic grade chloroform was used. Distilled water was used throughout the investigation.

A 2 M sulphuric acid (H<sub>2</sub>SO<sub>4</sub>) solution was prepared with concentrated H<sub>2</sub>SO<sub>4</sub> (S. D. Fine Chem., Mumbai, India, sp. gr. 1.84). 2 M sodium hydroxide (NaOH) was prepared in water. For the 0.1% (w/v) Orange II (ORG II) solution preparation, 500 mg of the dye (S. D. Fine Chem., Mumbai, India) were warmed with 200 mL water. Then, the solution was filtered through a Whatman No. 42 filter paper, and diluted to 500 mL with water. Tetrahydrofuran (THF), polyvinylchloride (PVC), dibutyl phthalate (DBP), dioctyl sebacate (DOS) and nitrophenyl octyl ether (NPOE) were purchased from S. D. Fine Chem.

Aqueous solutions of NaOH, potassium hydroxide, sodium carbonate, monopotassium phosphate, sodium acetate, sodium bicarbonate, acetic acid, potassium chloride, AgNO<sub>3</sub>, potassium nitrate, phosphoric acid, sodium nitrite, cobalt chloride, cadmium chloride, glucose, sucrose, oxalic acid, glycine, arginine, talc and urea (S. D. Fine Chem.) were prepared in bi-distilled water, by maintaining the required strength of 1 M each.

### ***Standard drug solution***

The pure Pharmaceutical grade HDH was offered by UCB Pharma Ltd, Mumbai, India. A standard 5.0 mmol/L<sup>-1</sup> HDH stock solution was prepared in water, and used for experimentation.

## **General procedures**

### ***Preparation of HDH-ORG II ion-associate***

10 mg of HDH were dissolved in 25 mL water, and 19 mg of ORG-II were added to the mixture. The contents were transferred to a clean separating funnel, and acidified with 1 mL of 2 M H<sub>2</sub>SO<sub>4</sub>, followed by the addition of 10 mL of

chloroform (CHCl<sub>3</sub>). The contents were shaken rigorously for 60 s, and the layers were allowed for equilibration. The organic layer was passed through anhydrous sodium sulphate, and put into a dry 100 mL beaker. The procedure was repeated twice with fresh 10 mL CHCl<sub>3</sub>. The combined extract was kept on a water bath, to evaporate the solvent, and the resultant solid HDH-ORG II ion-pair complex was collected.

#### ***Designing of the HDH-ORG II sensor***

10.0 mg HDH-ORG II ion pair complex, 0.16 g PVC and 0.1 g DBP were dissolved in 10 mL THF. The solution was transferred to a 4 cm wider Petri Dish, and dried under laboratory temperature, for 24 h. The thin membrane obtained with 3.7% (w/w) HDH-ORG II ion-associate, 59.2% (w/w) PVC and 0.37% (w/w) DBP was attached to one end of a plastic tube using THF, and dried at room temperature, for 24 h. 4 mL of 5.0 mmol/L<sup>-1</sup> HDH and 2.0 mmol/L<sup>-1</sup> KCl solutions were inserted into the tube. An Al wire was immersed into the tube, and connected to a potentiometer. The electrode was soaked in a 50 mmol/L<sup>-1</sup> HDH solution, and was used after 5 h.

#### **Procedure for bulk drug**

##### ***Construction of the calibration curve***

Suitable aliquots of a 5.0 mmol/L<sup>-1</sup> standard drug solution, equivalent to  $2.2 \times 10^{-5}$  -  $1.1 \times 10^{-3}$  mol/L<sup>-1</sup> of HDH, were measured and put into a series of 25 mL calibrated flasks. The volumes in each flask were raised to 30 mL with water, and the solutions pH in each flask was brought to the range from 2.5 to 5.21, in a 3 M NaOAc solution. The solutions were diluted up to the mark with water and well mixed, and the potentials were measured with the respective membrane sensor, using Ag/AgCl as reference electrode, at 25 °C.

The measured potentials calibration plot versus log[HDH] was constructed and used to find an unknown HDH concentration. The corresponding RE was also obtained and used to calculate HDH concentrations in the samples.

##### ***Procedure for tablet analysis***

20 Atarax 25 and Atarax 10 (UCB Pharma Ltd, Mumbai, India) tablets were roughly weighed for the required quantity, powdered and used for the analysis. An accurate amount of the powder equivalent to 50 mg HDH was transferred into a 100 mL calibrated flask, and 70 mL water were added to it. The mixture was shaken during 20 min, for proper dissolution, and the volume was brought up to the mark with distilled water. The solution was then filtered through Whatman no. 42 filter paper. A calculated volume of the tablet extract solution was diluted to obtain 5.0 mmol/L<sup>-1</sup> HDH, and a suitable aliquot was used for the assay that followed the procedure above described.

##### ***Interference study***

1 to 10.0 mL of the 5.0 mmol/L<sup>-1</sup> pure HDH solution, 10 mL water and 1 mL of the 1 mol L<sup>-1</sup> interferent solution were put into a series of 50 mL beakers. After the

solutions pH adjustment from 2.5 to 5.21, they were diluted to 25 mL with water, mixed for homogeneity, and their potentials were measured by the proposed sensor, relative to the Ag-AgCl reference electrode.

### Sensor $K_{HDH,I}$ determination

In a group of 50 mL beakers, 1 to 10 mL of a 5 mmol/L<sup>-1</sup> HDH solution were taken and mixed with 1.0 mL of 1.0 mol/L<sup>-1</sup> interferent. The solutions pH was adjusted in the range from 2.5 to 5.21 and, then, the volumes were raised with water to 25 mL, and homogenised. The new sets of solutions were prepared using the same procedure for different interferents. The potentials were recorded using the HDH-ORG II sensor.

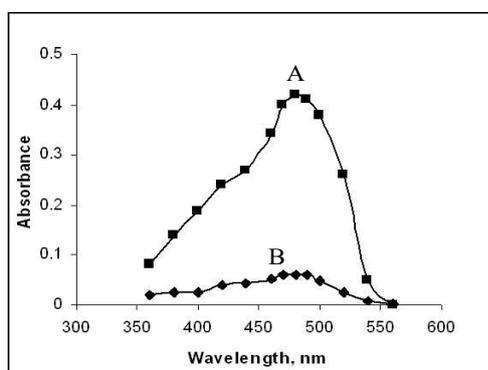
The graph of  $E_{Cell}$  (cell potentials) versus  $\log[HDH]$  was plotted, the point of intersection was located, and the  $K_{HDH,I}$  values for each interferent were calculated using the following formula [43]:

$$K_{HDH,I} = \frac{[HDH]_E}{[I]_E^{z_{HDH}/z_I}} = \frac{[HDH]_{Int}}{[I]_{add}^{z_{HDH}/z_I}}$$

where  $z_{HDH}$  is the cationic and anionic charges on HDH,  $z_I$  is the charge on the added interferent,  $[HDH]_E$  and  $[I]_E$  are the HDH and interferent concentrations, respectively, to generate indistinguishable  $E_{Cell}$  values,  $[HDH]_{Int}$  is the HDH concentration in the internal solution and  $[I]_{add}$  is the interferent concentration present in or added to the HDH solution.

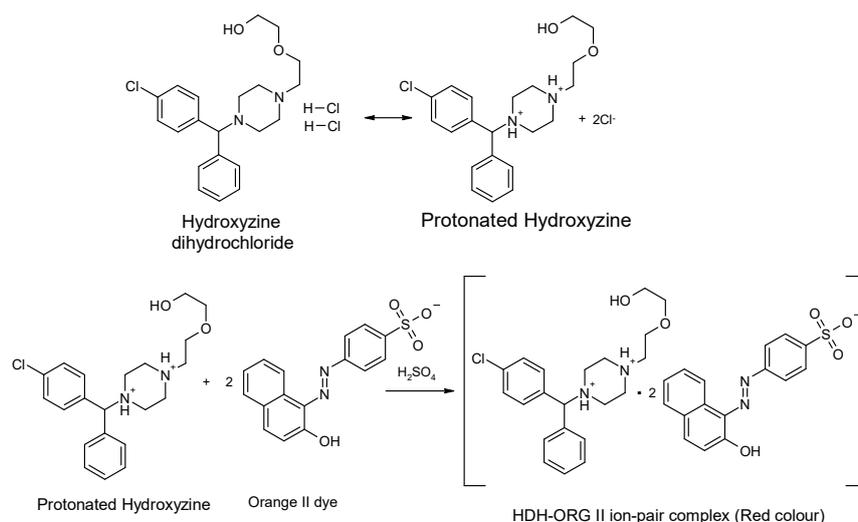
### Results and discussion

Due to the sulfonate anionic group present in the ORG II dye, there is a strong tendency to form the complex with the counter cations. In the membrane preparation, the dye was utilised as an ion-pair complexing agent with HDH in  $CHCl_3$ . Hydroxyzine, as a positively charged amino compound in an acidic medium, formed an extractable red coloured HDH, in  $CHCl_3$  [38], ion-pair complex with the ORG II anionic dye (1:2), which shows maximum absorption at 480 nm, as represented in Fig. 2.



**Figure 2.** Absorption spectra of A: ion-pair complex (9.0  $\mu\text{g/mL}^{-1}$  HDH with ORG II); and B: blank (ORG II without HDH).

The possible reaction pathway showing the formation of the ion-pair complex between the drug and the dye is shown in Scheme 1.



**Scheme 1.** Pathway of the HDH-ORG II ion-pair complex formation.

The resultant ion-pair complex was used in the membrane fabrication. The systematic representation of the electrochemical cell for HDH determination, using the membrane sensor, is shown below:



where Ag-AgCl<sub>IR</sub> and Ag-AgCl<sub>SR</sub> are the reference electrodes immersed into an internal (HDH<sub>i</sub>) and reference (HDH) sample solutions, respectively, and Membrane is a sensor devised with ORG II as an ion associating agent, DBP as plasticizer and PVC as matrix.

The relation between E<sub>Cell</sub> and [HDH]<sub>Sample</sub> is given by the Nernst equation [43]:

$$E_{\text{Cell}} = K + 0.059161 \log[\text{HDH}]_{\text{Sample}}$$

where K accounts for the reference electrode.

## Variables optimization

### Membrane components

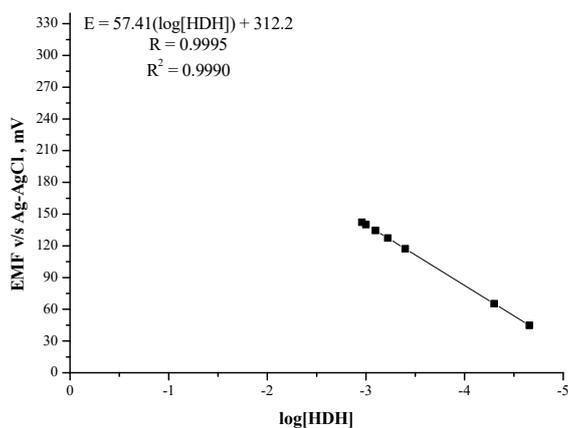
At the beginning, a series of trials, with different amounts of materials, such as ion exchangers, plasticizer and PVC, were executed to obtain the suitable membranes. The membranes obtained with different amounts of each of the above components were used for constructing ionic selective electrodes (ISEs), and their effective functioning for HDH sensing was checked by potentiometry. The results obtained for the membrane sensors, using 10 mg of ion-associate, 160 mg PVC and 100 mg DBP, were highly reliable. Moreover, the calibration line slope prepared with varying materials amounts, at other concentrations than those indicated above, was not acceptable for Nernstian behavior. 10 mL THF were found suitable for the complete materials dissolution. At larger volumes than 10 mL, no much variation was seen. The membranes became dried enough in 24 h, after being poured into a Petri Dish. Hence, the same standing evaporation time was fixed. Thus, the membranes were prepared by the optimised procedure.

### ***Plasticizer choice***

Different plasticizers, such as dibutyl sebacate (DBS), DBP, dioctyl phthalate (DOP) and NPOE, were used in different amounts, to obtain a suitable membrane. The Nernstian behaviour was seen in the membranes prepared with 100 mg DBP. The sensor showed a quick performance, with consistent potentials and ease of conditioning. However, the evaporation and response time were prolonged when the DBP amount was greater than 100 mg. So, 100 mg DBP is recommended to be used as plasticizer to design the HDH-ORG II sensor. Other plasticizers did not yield acceptable results in relation to calibration, Nernstian behaviour, quick response and consistent  $E_{Cell}$  values.

### ***Internal reference solution***

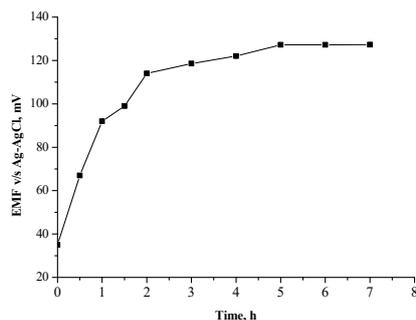
The calibration curve was prepared for HDH potentials and logarithmic concentrations data, using a HDH-ORG II sensor, with varied concentrations of internal HDH and KCl solutions, for obtaining linearity between the potentials, logarithmic concentrations and the Nernstian slope. The required result was obtained when a mixture of 5 mM HDH and 2 mM KCl solutions, as internal standard, was used. An acceptable correlation between potentials and  $\log[HDH]$  was observed, with a Nernstian slope of 57.41 mV/decade, for the proposed sensor. The calibration curve between potentials and  $\log[HDH]$  is shown in Fig. 3.



**Figure 3.** Calibration plots of  $E_{Cell}$  against  $\log[HDH]$ , for HDH concentrations in the range from  $2.2 \times 10^{-5}$  to  $1.1 \times 10^{-3}$  M, under optimum conditions, with the HDH-ORG II sensor.

### ***Soaking time***

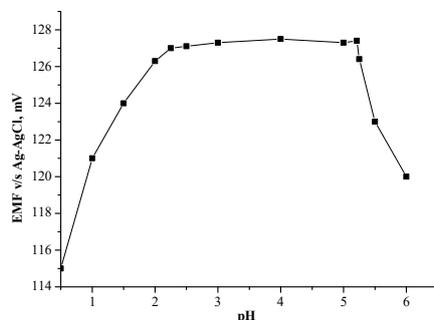
The sensors were soaked in a HDH standard solution, for different intervals of time to sufficiently activate their surfaces. The soaking was finalised by locating the time point from which the potential became steady and constant. The experimental results indicated the need of soaking the sensors for 5 h. Thus, this time was fixed as the duration required to make the membrane active surface ready for effective use at 25 °C. The fixed time factor and its variation, to measure the HDH sensors potential, are indicated in Fig. 4. This study also revealed and confirmed that the dried sensors may be placed in any closed opaque vessel, and used after conditioning, by soaking them in a standard HDH, for 5 h.



**Figure 4.** Effect of soaking/conditioning time on the potential of a  $6.6 \times 10^{-4}$  M HDH solution measured with a HDH-ORG II sensor.

### *Effect of pH*

The pH effect on the HDH  $E_{Cell}$  solutions using the proposed sensor, relative to the Ag-AgCl electrode, was evaluated. The HDH solutions  $E_{Cell}$  values, in the pH range from 0.5 to 10, were measured using the proposed electrochemical cell assembly. The solutions pH was brought to the required values by adding 1 M NaOAc or diluted  $NH_3$  solutions, and the potentials were recorded. The steady state was found to be attained in the pH range from 2.5 to 5.21. The potential values were lower and inconsistent out of the mentioned pH range. This pH effect on the HDH solutions  $E_{Cell}$  is showed in Fig. 5. According to the trend in the plot, the pH in the range from 2.5 to 5.21 was fixed as optimum for the HDH solutions potentials measurement using the HDH-ORG II sensor.



**Figure 5.** pH effect on the potential of a  $6.6 \times 10^{-4}$  M HDH solution measured with a HDH-ORG II sensor.

### *Response time*

The experimental response time of the conditioned sensor was evaluated. It was found to be 30 s for HDH solutions, after immersion in an analyte solution. Therefore, the conditioned sensor was immersed into HDH sample solutions, and the potentials were recorded after 30 s.

### *Sensor life time*

The sensor was assessed for its stable and uncompromised performance ability. The resultant Nernstian slope was free from deviations, while the sensor was used for, at least, 48 days. However, after this period, the Nernstian behaviour was disturbed. Therefore, 48 days of validity were declared for the devised sensor.

### **$K_{\text{HDH,I}}$ evaluation**

The chemical interaction between the membrane surface active sites and the analyte is responsible for the membrane potential ( $E_m$ ) generation. The membranes poor selectivity towards a single species is due to the chemical process on the signal.  $E_m$  is practically in proportion to the ions concentration capable of interacting at the membrane active sites. The generalized Nernst equation, that includes the contribution of an interferent, I, is convened as below:

$$E_{\text{Cell}} = K + \frac{0.05916}{Z_{\text{HDH}}} [A] ([\text{HDH}] + K_{\text{HDH,I}} [\text{I}]^{Z_{\text{HDH}}/Z_{\text{I}}}) \quad (1)$$

where K constant includes the reference electrode potential, the liquid junction potential and  $E_m$ . The charges on HDH and I are denoted by  $Z_{\text{HDH}}$  and  $Z_{\text{I}}$ , respectively. The  $K_{\text{HDH,I}}$  parameter is defined as:

$$K_{\text{HDH,I}} = \frac{[\text{HDH}]_E}{[\text{I}]_E^{Z_{\text{HDH}}/Z_{\text{I}}}} \quad (2)$$

where  $[\text{HDH}]_E$  and  $[\text{I}]_E$  are the HDH and interferent concentrations, respectively, to yield the same  $E_{\text{cell}}$  values. When  $K_{\text{HDH,I}}$  is equal to unity, the membrane responds for both HDH and I. A selective membrane must always possess a  $K_{\text{HDH,I}}$  value  $< 1.0$ .

Thus,  $K_{\text{HDH,I}}$  can be evaluated using the solutions potentials with I and  $[\text{I}]_{\text{add}}$  known but constant amounts, and varying HDH concentrations. Two distinct linear regions in the  $E_{\text{cell}}$  plot against the analyte concentration logarithmic values gave significant data for the  $K_{\text{HDH,I}}$  value assessment. When  $[\text{HDH}]_E$  is significantly  $> K_{\text{HDH,I}} [\text{I}]_{\text{add}}$ ,  $E_{\text{cell}}$  is a linear function of the log  $[A]$ . However, if  $K_{\text{HDH,I}} [\text{I}]_{\text{add}}$  is significantly  $> [A]$ ,  $E_{\text{cell}}$  remains constant. The analyte and interferent concentrations, at the two linear regions intersection, were used to calculate  $K_{\text{HDH,I}}$  for HDH, with different Is (Table 1). The sensor was tested with various inorganic and organic compounds, anions and cations. For the tests, a 1 M solution of each interferent was spiked into a pre-analysed HDH solution [44, 45]. The results show that the developed sensor was free from the interfering species action during the HDH analysis, as shown by the  $K_{\text{HDH,I}}$  values (Table 1).

**Table 1.** Sensor  $K_{\text{HDH,I}}$  for different interferents.

Interferent	$K_{\text{HDH,I}}$
-Na <sup>+</sup>	0.235
K <sup>+</sup>	0.213
H <sup>+</sup>	0.123
Ca <sup>2+</sup>	0.342
Zn <sup>2+</sup>	0.234
Glycine	0.090
Ammonium oxalate	0.063
Formic acid	0.089
Citric acid	0.127
Tartaric acid	0.342
Benzoic acid	0.211
Salicylic acid	0.132
Phthalic acid	0.187
Oxalic acid	0.127
Boric acid	0.389
Talc	0.474

\*Average of 5 determinations.

Therefore, the developed sensor is the most suitable electrode for HDH analysis, even with such interferents.

### Sensor validation

The sensor has been validated for various parameters, according to IUPAC [44, 45] and the International Council for Harmonization of Technical Requirements for Pharmaceuticals for Human Use (ICH) guidelines [46]. The results are presented in the following sections.

#### *Linearity of calibration curve, regression data and performance characteristics*

HDH measured potentials and concentrations were in very good correlation with linearity, as shown in the calibration curves in Fig. 3. The linearity obeying the Nernstian behavior has been reflected by the slope of  $57.41 \pm 1.04$  mV/decade. The curve fitting equation from the regression data is obtained as  $Y = 57.41 X + 312.2$ . The RC value, with very good linearity between the measured potentials and  $\log[\text{HDH}]$ , has also been calculated. All these data, along with the LOD, were calculated by following IUPAC's recommended procedures [44, 45]. The values measured from the intersection of the calibration line extrapolated linear portions and other parameters for the sensor performance are given in Table 2.

**Table 2.** HDH-ORG II membrane sensor performance features.

Parameter	Values
Linear range, M	$2.2 \times 10^{-5}$ to $1.1 \times 10^{-3}$
Limit of detection	$4.5 \times 10^{-6}$
Slope (m),	57.41
Intercept (b), mV	312.2
Correlation	0.9995
R <sup>2</sup>	0.9990
pH (optimum)	2.50 – 5.21
Life time, days	48

#### *Evaluation of precision and accuracy*

For studying precision with intra-day variations, seven replicates, each of 0.4, 0.6 and 0.8 mM HDD, were prepared by diluting a 0.005 M HDH standard solution with distilled water, and analysed. The %RSDs for each solution were calculated (Table 3). For inter-day deviations, the analysis of five replicates, each of 0.4, 0.6 and 0.8 mM HDH, were analysed in three days.

**Table 3.** Results of precision and accuracy of the HDH-ORG II sensor.

HDH taken mM	Intra-day variations			Inter-day variations		
	HDH found* mM	%RSD	%RE	HDH found <sup>§</sup> mM	%RSD	%RE
0.40	0.39	4.56	2.50	0.41	2.66	2.50
0.60	0.61	4.81	1.67	0.62	3.11	3.33
0.80	0.78	4.32	2.50	0.84	1.86	5.00

\*Mean value of seven measurements; <sup>§</sup>Mean value of five measurements.

The calculations were also undertaken to obtain the %RE (relative error) inter- and intra-day analyses values, and used as an accuracy metric. The %RSD values, from

1.86 to 4.81, and the %RE values, from 1.67 to 5.0, indicate the accurate and precise functioning of the ion electrode for HDH determination.

### **Robustness and ruggedness**

The sensor robustness was assessed by purposely changing the optimum operating temperature by 2 °C, in the analysis of 0.4, 0.6 and 0.8 mM HDH solutions. The %RSD values for HDH, at 23, 25 and 27 °C, were found to be in the range from 4.34 to 5.67 (Table 4). The proposed sensor was used for analysing HDH by analytes. The instrumental variations were also evaluated by three different potentiometers. The %RSD values variation of <5.7 confirmed the sensor robust and rugged behaviour.

**Table 4.** HDH-ORG II sensor robustness and ruggedness results (expressed in %RSD).

HDH concentration mM	%RSD values for varied parameters		
	Robustness (by varying T)	Ruggedness	
		Inter- analytes	Inter-potentiometric
0.40	4.34	3.34	4.26
0.60	4.65	3.10	3.82
0.80	5.67	4.67	3.99

### **Applicability of the HDH-ORG II sensor for HDH tablets analysis**

The analyses were done by taking five replicates of tablet extracts containing 0.6, 0.8 and 1.1 mM HDH, and using the designed and validated sensor. The found HDH amount and recovery values (%), with the RSD values, were calculated. The mean recovery and RSD values were compared with those of the official United States Pharmacopeia (USP) method [18], which describes an HPLC procedure with uv-detection at 232 nm, to determine HDH in tablets, by the Student-t and F-tests. The tablet extract HDH analytical recovery values obtained by the proposed and official methods,  $t_{Cal}$  and  $F_{Cal}$  (calculated), are all presented in Table 5. The  $t_{Cal}$  and  $F_{Cal}$  values were lower than 2.77  $t_{tab}$  and 6.39  $F_{tab}$  (tabulated), respectively, at a 95% confidence level, for four degrees of freedom. Hence, the sensors accurate and precise results inference was statistically given.

**Table 5.** HDH tablets analysis results using the proposed sensor, and statistical comparison with those of the official USP method.

Tablets analyzed	HDH tablet mg	Found*	
		%Label claim±SD	
		USP method	Proposed method using HDH-ORG II sensor
Atarax 25 <sup>a</sup>	25	97.26±1.18	98.16±0.69 t = 1.52 F = 2.92
Atarax 10 <sup>a</sup>	10	98.34±0.89	100.01±1.24 t = 2.48 F = 1.94

\*Mean value of 5 determinations. (The  $t_{tab}$  value is at the 95% confidence level, and, for four degrees of freedom, is 2.77). (The  $F_{tab}$  value is at the 95% confidence level, and, for four degrees of freedom, is 6.39).

<sup>a</sup>Marketed by UCB Pharma Ltd, Mumbai, India.

### **Recovery study**

The proposed HDH-ORG II sensor was used for recovery experiments, by following the standard addition procedure. Three different levels of pure HDH solutions spiked into a pre-analyzed tablet extract were obtained, and the potentials were recorded by potentiometry, using the HDH-ORG II sensor. Into five replicates of 2 mL tablet extract of 5 mM HDH, 1, 2 and 3 mL of a 5 mM HDH standard solution were spiked. The pH was adjusted to an optimum value and, after adding water up to 25 mL and mixing, the potentials were measured. The total HDH amounts were calculated in each case, and, finally, the recovery values were computed. The recovered HDH values (Table 6) were from 95 to 103.3, which highlighted the assay procedures accuracy using the proposed sensor.

**Table 6.** Accuracy assessment results in the recovery study by the standard-addition procedure.

HDH tablet extract mM	Pure added HDH mM	Total HDH found mM	%HDH recovered*	%RSD
2.00	1.00	2.95	95.00	2.78
2.00	2.00	4.05	102.50	3.14
2.00	3.00	5.10	103.30	4.22

\*Mean value of three measurements.

### **Conclusions**

A simple and cost effective PVC mediated ion selective electrode has been newly devised, for HDH potentiometric assay. It uses a sensor with ORG II as ion-exchanger, DBP as plasticizer, and PVC as matrix component, for HDH selective and rapid determination. The method does not pose any environmental hazard and, thus, it has a highly green approach. A simple potentiometer and eco-friendly analytes were required, which has made the method economic. The reported literature has described many analytical procedures for HDH determination, but they have disadvantages such as the need of sophisticated instruments, highly skilful operators, and costly and toxic solvents. Thus, this green sensor is useful within a wide linear range, with Nernstian response and a low LOD. The statistical tests confirmed the procedure accessibility for HDH assay in pure and tablet forms, with excellent recoveries, and agreement with the official USP method results. The excipients in tablets did not show any interference, and, hence, the method is extremely selective. Therefore, the HDH-ORG II sensor can be used as an analytical tool to determine HDH in quality control laboratories, on a routine basis.

### **Acknowledgments**

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### **Authors' contributions**

**Rajendra Prasad N.:** conceived and designed the analysis; contributed with data or analysis tools; wrote the paper. **Siddaraju C.:** collected the data; contributed

with data or analysis tools; performed the analysis; assisted in the paper preparation and revision.

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